
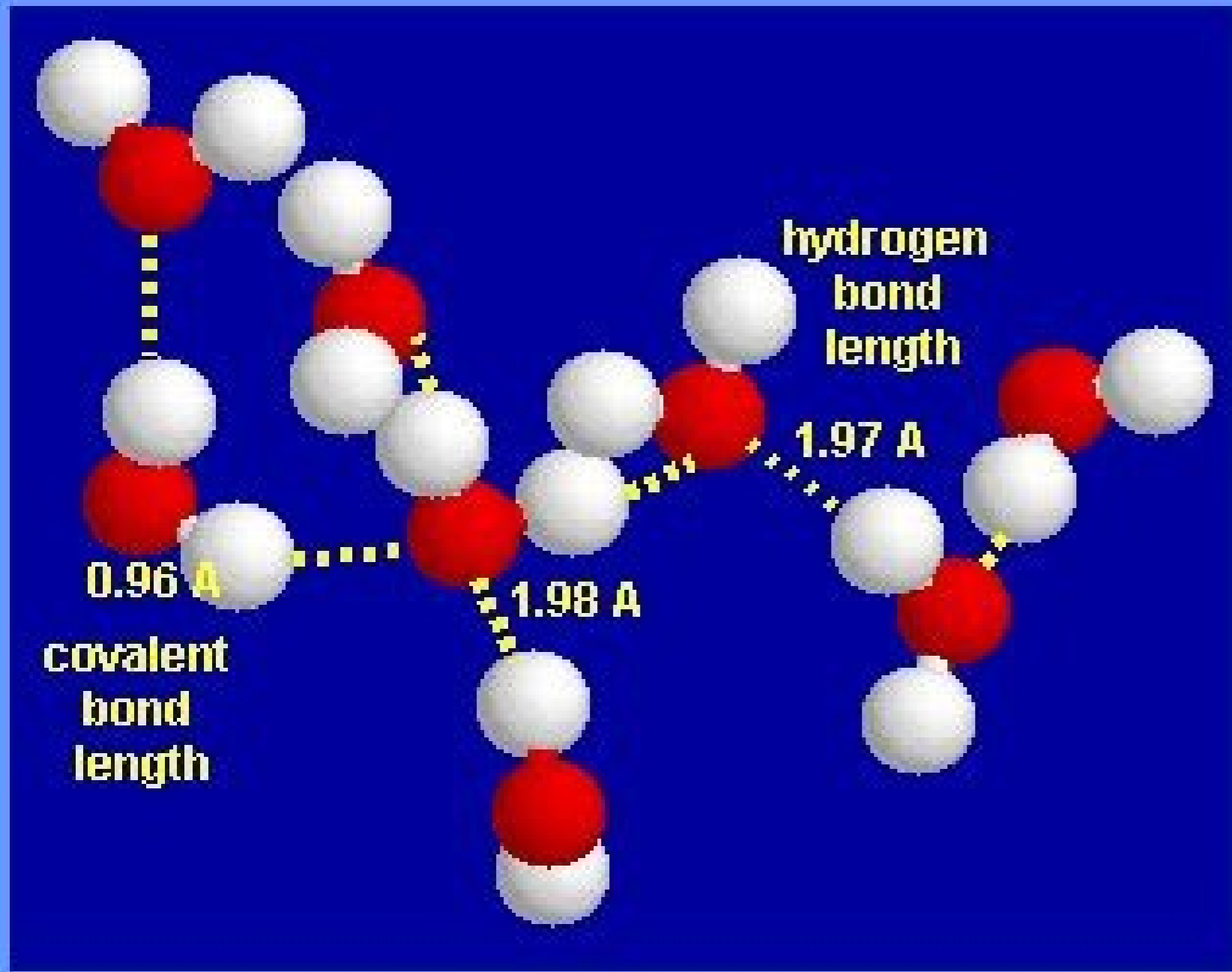


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# Hydrogen Bonds



C. Ophardt, c. 2003

Polarity and intermolecular forces gizmo worksheet answers. Student exploration polarity and intermolecular forces worksheet answers. Molecular polarity and intermolecular forces worksheet. Polarity and intermolecular forces worksheet answers.

Polar molecules have dipole-dipole forces. SO<sub>2</sub> Dipole-dipole forces SO<sub>2</sub> is a folded polar molecule. Intermolecular force Key Worksheet 1. Dipole-dipole forces CH<sub>2</sub>Cl<sub>2</sub> The strongest intermolecular force in a polar molecule that cannot form hydrogen bonds<sup>1</sup> is the dipole-dipole force e. Identify the strongest intermolecular force operating in the condensed phases of the following substances. Then, identify the strongest intermolecular force present in pure samples of the following substances: SO<sub>2</sub> H<sub>2</sub>O CH<sub>2</sub>Cl<sub>2</sub> dipole-dipole forces the hydrogen bonds dipole-dipole SCO<sub>2</sub> PCI<sub>3</sub> Dipole-dipole forces SO<sub>3</sub> dipole-dipole forces Dispersion forces<sup>2</sup> n London 2. Dipole-dipole forces CH<sub>3</sub>-O-CH<sub>3</sub> Hydr<sup>3</sup> gen is not attached to oxygen, so this mol<sup>4</sup> cannot form hydrogen<sup>5</sup> bonds. The strongest intermolecular force in a polar molecule is the dipole-dipole force d. Explain fully how<sup>6</sup> determines it. They also have dispersal forces<sup>7</sup> in Londoners, but the dipole-dipole forces are stronger. b. It's polar, so it will have dipole-dipole forces. a. Thank you for your participation<sup>8</sup> {{getToolbarWorksheetName()}} has been added to spreadsheets. Worksheet added to worksheets Don't forget to leave a comment. Please leave a comment. These are the strongest intermolecular forces. Molecules that have hydrogen attached<sup>9</sup> an O, N, or F can form hydrogen bonds<sup>10</sup>n. g. Print {{ws\_solutions.user.firstname}} {{ws\_solutions.user.lastname}} responds to {{ws\_solutions.user.username}} replies Assignment<sup>11</sup> closed Start creating - "Free! The C-O bond is polar, so the molecule is polar. Dispersion forces<sup>12</sup> n Cl<sub>2</sub> London The Cl-Cl link is non-polar so the mole is non-polar. Remove the following substances. The non-polar molecules<sup>13</sup> have forces of dispersion<sup>14</sup> n That act on the substance. c. c.

Balancing Equations. The chemical equation described in section 4.1 is balanced, meaning that equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides. This is a requirement the equation must satisfy to be consistent with the law of conservation of matter. Example 2. Computing Formula Mass for an Ionic Compound Aluminum sulfate, Al<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>, is an ionic compound that is used in the manufacture of paper and in various water purification processes. What is the formula mass (amu) of this compound? Solution The formula for this compound indicates it contains Al<sup>3+</sup> and SO<sub>4</sub><sup>2-</sup> ions combined in a 2:3 ratio. For purposes of ...

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